Standard Reduction Potentials

Ox + ne^{1-}		**Red**	$\mathcal{E}^\circ$ **(V)**
$K^{1+}(aq) + e^{1-}$	$\leftrightharpoons$	$K(s)$	-2.92
$Na^{1+}(aq) + e^{1-}$	$\leftrightharpoons$	$Na(s)$	-2.71
$Mg^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Mg(s)$	-2.36
$Al^{3+}(aq) + 3\ e^{1-}$	$\leftrightharpoons$	$Al(s)$	-1.66
$2\ H_2O + 2\ e^{1-}$	$\leftrightharpoons$	$H_2(g) + 2\ OH^{1-}(aq)$	-0.83
$Zn^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Zn(s)$	-0.76
$Cr^{3+}(aq) + 3\ e^{1-}$	$\leftrightharpoons$	$Cr(s)$	-0.74
$Fe^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Fe(s)$	-0.44
$2\ H_2O + 2\ e^{1-}$	$\leftrightharpoons$	$H_2(g) + 2\ OH^{1-}(aq)$	-0.41^*
$PbSO_4(s) + 2\ e^{1-}$	$\leftrightharpoons$	$Pb(s) + SO_4^{2-}(aq)$	-0.36
$Ni^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Ni(s)$	-0.23
$Sn^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Sn(s)$	-0.14
$Pb^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Pb(s)$	-0.13
$2\ H^{1+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$H_2(g)$	0.00
$Cu^{2+}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$Cu(s)$	$+0.34$
$O_2(g) + 2\ H_2O + 4\ e^{1-}$	$\leftrightharpoons$	$4\ OH^{1-}(aq)$	$+0.40$
$I_2(s) + 2\ e^{1-}$	$\leftrightharpoons$	$2\ I^{1-}(aq)$	$+0.54$
$Fe^{3+}(aq) + e^{1-}$	$\leftrightharpoons$	$Fe^{2+}(aq)$	$+0.77$
$Ag^{1+}(aq) + e^{1-}$	$\leftrightharpoons$	$Ag(s)$	$+0.80$
$O_2(g) + 4\ H^{1+}(aq) + 4\ e^{1-}$	$\leftrightharpoons$	$2\ H_2O$	$+0.82^*$
$NO_3^{1-}(aq) + 4\ H^{1+}(aq) + 3\ e^{1-}$	$\leftrightharpoons$	$NO(g) + 2\ H_2O$	$+0.96$
$Br_2(l) + 2\ e^{1-}$	$\leftrightharpoons$	$2\ Br^{1-}(aq)$	$+1.09$
$O_2(g) + 4\ H^{1+}(aq) + 4\ e^{1-}$	$\leftrightharpoons$	$2\ H_2O$	$+1.23$
$Cr_2O_7^{2-}(aq) + 14\ H^{1+}(aq) + 6\ e^{1-}$	$\leftrightharpoons$	$2\ Cr^{3+}(aq) + 7\ H_2O$	$+1.33$
$Cl_2(g) + 2\ e^{1-}$	$\leftrightharpoons$	$2\ Cl^{1-}(aq)$	$+1.36$
$MnO_4^{1-}(aq) + 8\ H^{1+}(aq) + 5\ e^{1-}$	$\leftrightharpoons$	$Mn^{2+}(aq) + 4\ H_2O$	$+1.51$
$PbO_2(s) + 4\ H^{1+}(aq) + SO_4^{2-}(aq) + 2\ e^{1-}$	$\leftrightharpoons$	$PbSO_4(s) + 2\ H_2O$	$+1.69$
$F_2(g) + 2\ e^{1-}$	$\leftrightharpoons$	$2\ F^{1-}(aq)$	$+2.87$

Standard Reduction Potentials at 25 °C in volts

* The half-cell potentials for the O_2/H_2O and the H_2O/H_2 systems are for pure water and are not for standard reduction potentials where $[OH^{1-}]$ and $[H^{1+}] = 1.0\ M$.